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I, LEANNE MYNOTT, MANAGER EXAMINATION SUPPORT AND SALES hereby certify that annexed is a true copy of the Provisional specification in connection with Application No. 2003906041 for a patent by OXIANA RESOURCES NL as filed on 03 November 2003.

I further certify that the above application is now proceeding in the name of OXIANA LIMITED pursuant to the provisions of Section 113 of the Patents Act 1990.



WITNESS my hand this
Sixteenth day of November 2004

A handwritten signature in black ink, appearing to be 'L. Mynott'.

LEANNE MYNOTT
MANAGER EXAMINATION SUPPORT
AND SALES

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Leach Extraction Method

Field of the Invention

This invention relates to the leach extraction of minerals. This invention in particular
 5 relates to the high temperature autoclave leaching of copper – iron – sulfur
 concentrates for the purpose of extracting copper and producing a strong acid and
 ferric sulfate solution for associated copper leaching operations, and for illustrative
 purposes the invention is described hereinafter with reference to this application.
 However, it is envisaged that this invention may find application in other mineral
 10 extractions.

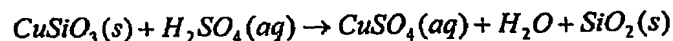
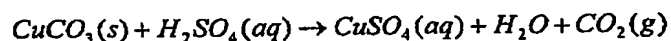
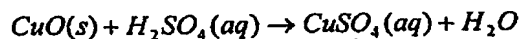
Background

The atmospheric leaching of copper ores can be carried out using a variety of
 contactor types. Copper ores can be leached on heaps or dumps, in vats or in stirred
 15 tanks. The choice of the type of contactor will vary with the type and grade of the ore
 and its leach characteristics as well as local economic, environmental, topological and
 hydrogeological character.

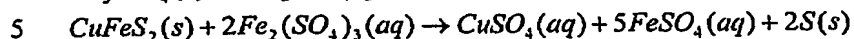
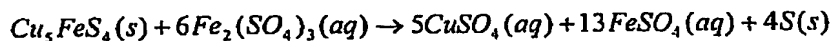
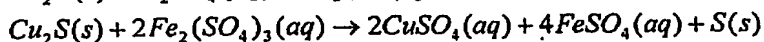
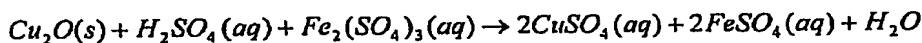
The leaching of copper minerals generally proceeds by direct acid attack for simple
 20 oxides, carbonates and silicates while acid – oxidizing conditions are necessary to
 leach reduced copper species or copper sulfide minerals. Examples of the chemistry
 of leaching of different copper minerals are shown below.

Acid leaching reactions:

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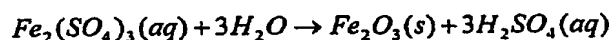
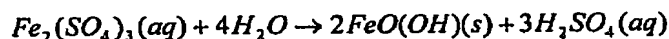
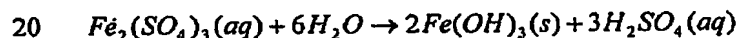
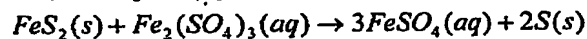
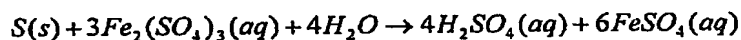
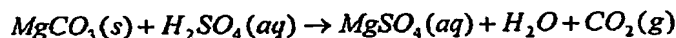
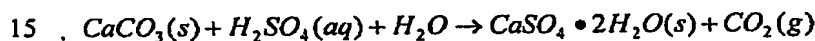


30 Acid – oxidizing leaching reactions using ferric sulfate as an oxidant:



The rate and extent of copper leaching obtained in this series of reactions is directly linked to the availability of acid or acid and ferric ion.

- 10 Acid and ferric ion can also be consumed by "side" reactions with barren (of copper) minerals such as carbonates or other sulfides or sulfur. Ferric ion can also be consumed by hydrolysis to form precipitates such as ferric hydroxide, goethite, hematite and jarosite.

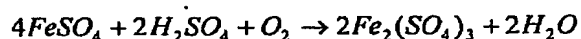


- 25 Note that the interaction of these reactions can be complex with some reactions consuming acid and some producing acid.

In summary, in industrial leaching operations for copper, acid and ferric ion are required in sufficient amounts to complete the leaching reactions to maximize the

- 30 extraction of copper.

In many leaching operations, acid is purchased and added as a reagent to the leach solution. Generally, ferric ion is produced *in situ* in leaching systems, for example, by oxidation of ferrous sulfate to ferric sulfate. In lower temperature systems, the oxidation of ferrous to ferric is catalyzed by natural bacterial action in the presence of oxygen from air. At higher temperature, superatmospheric-pressure oxygen is often supplied to accelerate the rate of oxidation.

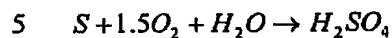
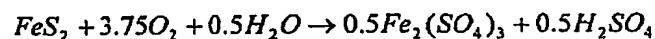
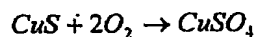
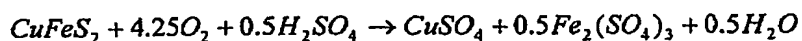


10 An example of such a process is described in International Patent Publication WO 00/06784. This document describes a process suitable for a high-pyrites content ore. High pyrites-content ores have sulfur content sufficient to regenerate acid *in situ*, making this process suitable for these types of ores. In many locations where copper is leached, the cost to purchase and transport acid for leaching of copper is
15 prohibitively expensive, where the ore type and grade is insufficient for maintenance of sulfuric acid levels *in situ*, that is, without making up the acid levels with exogenous sulfuric acid. Hitherto, it has not been possible to conveniently make rich ferric sulfate solutions for copper leaching at the same time as acid is formed in the autoclave, for low pyrites ores.

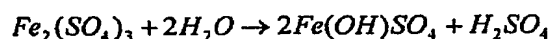
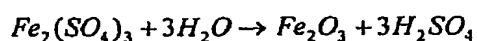
20 The reason for this is that the production of rich ferric sulfate and sulfuric acid is favored by higher autoclave temperatures and oxygen pressures. However, these same conditions favor side reactions that generate precipitated species in lieu of acid or consume acid resulting in the requirement for acid make-up. Also, in the case of
25 high-pyrites containing ores, side reactions result in the precipitation of elemental sulfur, which creates processing difficulties at elevated temperatures at which viscous allotropes of sulfur form.

During autoclave pressure oxidation leaching of copper and iron sulfides and sulfur, a
30 number of reactions will occur. These can be classified as oxidation reactions and precipitation reactions. For example, at high temperature (+150°C):

Oxidation (all sulfide sulfur oxidizes to sulfate)



Iron Precipitation Reactions

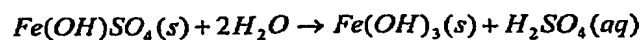


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The oxidation reactions increase the solution content of dissolved copper, ferric ion and acid. The precipitation reactions precipitate iron as either hematite or basic ferric sulfate. In the first precipitation reaction, all sulfate stays in solution and acid is produced in significant amounts (three acid formed for each ferric sulfate). The

15 second precipitation reaction forms basic ferric sulfate. In this reaction, the formation of acid is severely curtailed as the basic ferric sulfate contains two of the three available sulfates.

The formation of hematite as the iron precipitation product results in a high strength
20 acid solution from the autoclave process. The formation of basic ferric sulfate has generally been viewed as undesirable as the basic ferric sulfate dramatically reduces the strength of acid and ferric sulfate in the autoclave discharge solution. From an environmental perspective, basic ferric sulfates are undesirable as basic ferric sulfates will gradually decompose in tailings impoundments, resulting in slow release of acid
25 and ferric ion. This decomposition of basic ferric sulfate can acidify a tailing and result in acid mobilization of any contaminants in the tailings solids.



30

Unfortunately, the formation of basic ferric sulfates is favoured by (1) higher temperature and (2) increasing concentration of dissolved salts. For example, as the magnesium sulfate level is increased in solution, the "break" point indicating the onset of basic ferric sulfate precipitation advances to lower free acid concentrations.

5

To summarize, under certain operating conditions for pressure oxidation of copper/iron/sulfur containing ores, concentrates or residues, the autoclave leach solution will contain dissolved copper and ferric sulfate salts and sulfuric acid while the residues will contain hematite and basic ferric sulfate. The presence of basic ferric sulfate reduces (1) the available acid in the autoclave solution (acid formation by iron precipitation is attenuated) and (2) the available ferric sulfate in the autoclave solution. In addition, the presence of basic ferric sulfate will render the autoclave residue environmentally unstable. For these reasons, to avoid basic ferric sulfate formation, autoclave conditions are controlled by (1) lowering the operating temperature, (2) reducing the pulp density (solid to liquid ratio) within the limits of an overall heat balance and (3) leaching in water rather than leaching in available sulfate containing solutions. All of these control strategies are undesirable as they result in increased costs or processing complexity. For example, at lower temperature, all oxidation reactions are slower and therefore a longer autoclave oxidation time is required. This would necessarily require a larger autoclave for treatment at the lower temperature. Similarly the reduction of pulp density results in movement of more water and less solid through the autoclave, again increasing the size of the autoclave. Finally, leaching in water rather than leaching in available sulfate containing solutions may unreasonably constrain the operation of a commercial autoclave facility by disrupting the overall site "water balance".

It is one purpose of embodiments of the present invention to provide a ready source of both sulfuric acid and ferric ion from the autoclave oxidation of copper/iron/sulfur containing feed material. It is a purpose of certain embodiments of the present invention to provide a source of ferric ion for the production of a strong oxidizing solution suitable for oxidizing other minerals such as zinc sulfides, uranium oxides, nickel and cobalt sulfides, and many others.

30

Detailed Description of the Invention

5 In one aspect the present invention resides broadly in a method for leach extraction of mineral bearing ores and concentrates including the step of leaching said mineral with an aqueous stream containing ferric ions and sulfuric acid in the presence of oxygen, at least part of said aqueous stream comprising a solution formed by reaction of basic ferric sulfate with excess sulfuric acid.

10 The basic ferric sulfate may be produced by any suitable means. For example, the basic ferric sulfate may be produced as a by-product of minerals processing. Alternatively the basic ferric sulfate may be produced *de novo* in a process such as the autoclave oxidation of iron-bearing pyrite mineral slurry at elevated temperatures and under superatmospheric-pressure oxygen. Typically production of basic ferric sulfate
15 precipitate is favored by temperatures above 150°C, preferably above 200°C, and oxygen overpressures in excess of 4 bar, typically about 6 bar.

The iron-bearing pyrite mineral may be native pyrite ore or may be produced as a by-product of mixed-ore processing. For example, the iron-bearing pyrite mineral may
20 be obtained from the solids residue of processing copper/iron/sulfur bearing ores such as chalcocites. In particular embodiments of the present invention the iron-bearing pyrites mineral forms a part of the solids separated from an atmospheric ferric/acid leach of aqueous slurries of chalcocite ores. The separated solids may be concentrated by flotation to produce a pyrites concentrate and tailings stream, and this
25 is particularly preferred in the case of ores having relatively low pyrite content.

The basic ferric sulfate leaching step may comprise the addition of acidified water to the solid basic ferric sulfate and aging the slurry for a time and at a temperature sufficient for the supernatant to charge with ferric ions. Alternatively, the basic ferric
30 sulfate leaching step may simply comprise the releaching of the basic ferric sulfate slurry from the autoclave of its production, wherein the temperature of the slurry is reduced to a temperature that favors the releaching of the basic ferric sulfate to form a

supernatant of strong ferric and acid content. The temperature and time of residence must be determined empirically since it will depend on the source of the basic ferric sulfate and the releaching solution. However, as a guide it has been determined that autoclave-formed precipitate containing basic ferric sulfate may be leached by
5 acidified solution or fresh water at temperatures below about 100°C, and typically about 95°C, and autoclave slurries may be cooled to at least below 100°C to 120°C depending on chemical environment and preferably to about 85°C for releaching.

In particular embodiments of the present invention, a primary ore stream may be
10 directly autoclaved to produce a first winnable raffinate which may be processed to strip acid, and a solids slurry containing basic ferric sulfate that may be admitted, either directly or after separate releaching, to an atmospheric leach (with concurrent releaching of basic ferric sulfate) of a second ore slurry stream, with or without the stripped acid as necessary. This achieves the primary object of producing basic ferric
15 sulfate in the autoclave followed by an atmospheric redissolution of basic ferric sulfate in an acidic solution to produce an acidic ferric sulfate solution for use in leaching of copper minerals.

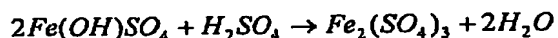
In another aspect this invention resides broadly in a method for leach extraction of
20 relatively low pyrites-containing chalcocite ores including the steps of:

- (a) atmospheric leaching of said ore with a contactor solution including ferric ion and sulfuric acid;
- (b) passing the pregnant leach solution of step (a) to copper winning;
- (c) autoclaving an aqueous slurry of the pyrites-containing solids from
25 step (a) in the presence of oxygen at a temperature and a time to produce basic ferric sulfate precipitate as a component of the autoclave discharge solids;
- (d) reacting said basic ferric sulfate within the autoclave discharge solids with aqueous sulfuric acid to form a solution containing
30 ferric ion and sulfuric acid; and
- (e) recycling said solution to step (a).

In preferred embodiments, the invention involves the precipitation of basic ferric sulfate at high temperature which may be conveniently redissolved by lowering the temperature of the autoclave discharge slurry, resulting in production of a rich ferric sulfate solution for leaching reduced copper minerals. In this case the aqueous acid
5 required for redissolution is formed as a consequence of the autoclaving process. The lowering of the temperature may be a consequence of pumping the acidified basic ferric sulfate slurry directly to the atmospheric leach step, where mixing with the incoming ore slurry reduces the temperature to permit redissolution. Alternatively the basic ferric sulfate precipitate may be filtered off ahead of redissolution with aqueous
10 acid to form an acid ferric sulfate solution for leaching in step (a).

The aqueous slurry from step (a) may be passed directly to autoclaving if the pyrite content is sufficiently high. In other cases the pyrite content may be enhanced by producing a concentrate by any known means such as flotation or the like.

15 In this work a method was developed to overcome the problems associated with basic ferric sulfate formation. First, conditions were set for pressure oxidation where formation of basic ferric sulfate occurred. The basic ferric sulfate in the autoclave discharge solids was then selectively re-dissolved at temperatures up to the boiling -
20 point by allowing acid produced in the autoclave to attack the basic ferric sulfate formed at elevated temperature. The chemistry of re-dissolution is shown below:



25 The re-dissolution of basic ferric sulfate allows the autoclave design engineer or operator the freedom to purposely form basic ferric sulfate in the pressure oxidation of copper - iron - sulfur feed materials. By extension, this also allows the designer/operator the ability to increase temperature to accelerate the rate of oxidation, to increase pulp density within the limits of the overall heat balance so as to
30 minimize autoclave size and finally to perform the pressure oxidation in whatever solution is most appropriate in view of the overall site water balance.

A further benefit of this re-dissolution of basic ferric sulfate dramatically reduces the environmental issues associated with basic ferric sulfate decomposition in tailings disposal of autoclave discharge solids.

- 5 Finally, the "re-dissolution" solution from acid attack of the basic ferric sulfate is now rich in both acid and ferric sulfate content and is an ideal leach solution for further leaching of copper (or other metals in other applications).

- 10 Further, the re-dissolution reaction can be affected by for example, holding the autoclave discharge slurry at temperatures of up to the boiling point under atmospheric pressure, or at higher temperatures with overpressure. Once the basic ferric sulfate has been largely converted to soluble ferric sulfate, the slurry can be directed toward a leaching process for copper in which acid and ferric sulfate are required. Alternately, the slurry can be treated to separate a solid residue (barren of
15 basic ferric sulfate) and an acidic ferric sulfate solution that can then be directed to a leaching process (without forwarding the solid residue at the same time).

- In an alternate embodiment, the autoclave discharge slurry can be mixed with, for example, a copper - iron - sulfide raw material. The advantage of this embodiment is
20 that the re-dissolution of basic ferric sulfate and the leaching of the copper - iron - sulfide raw material with the acidic ferric sulfate solution can proceed simultaneously.

- The invention will be further described with reference to the following examples and the Figures, wherein:
25

- FIG. 1 is a flowsheet of a process in accordance with the present invention;
FIG. 2 is a flowsheet of an alternative process in accordance with the present invention;
FIG. 3 is a flowsheet of a yet further alternative process in accordance with the
30 present invention; and
FIG. 4 is a graphic representation of results of an atmospheric leach of residue suitable for use in processes in accordance with the present invention.

These three processes are shown schematically below in a series of flowsheets.

The following examples illustrate the method of the invention.

5

Example 1

A 200 gm sample of solid containing 6.91% Cu, 29.1% Fe, 43.2% S(total), 33.2% S(sulfide), 9.03% S(elemental) and 0.95% S(sulfate) at P80 of 108 um particle size was mixed with 2 L of solution containing 4 g/L Cu, 12.98 g/L Fe (total), 12.0 g/L Fe(+3), 0.98 g/L Fe(+2) and 23.39 g/L H₂SO₄, was oxidized at 220°C for 25 minutes with 6 bar oxygen overpressure.

At the end of the test, the autoclave was quickly cooled, emptied and subjected to filtration and washing to separate solids and liquids. The leach solution and leach solid were collected, measured and analyzed.

The leach solution recovered from this test was 1990 mL with 8.6 g/L Cu, 13.8 g/L Fe(total), 1.1 g/L Fe(+2), 12.7 g/L Fe(+3) and 73.3 g/L H₂SO₄.

20

The leach solid residue weighed 197.7 gm and analyzed 0.16% Cu, 27.7% Fe, 17.8% S(total), 2.7 S(sulfide), 1.1% S(elemental) and 14.0 % S(sulfate).

Inspection of these results shows that copper extraction was ~ 96.7% and that most of the sulfur was oxidized to sulfate. The iron content of the leach solution was largely unchanged (13.8 g/L total versus an initial 12.98 g/L) while the acid level increased significantly. Much of the sulfate formed by oxidation coprecipitated with the iron in the autoclave residue as basic ferric sulfate. This is the reason for 14.0% S(Sulfate) in the autoclave discharge residue.

30

This example illustrates the ready formation of basic ferric sulfate in the autoclave 220°C

Example 2.

Test No	AC409
Description 3	220°C leach with atmospheric post leach to releach 'basic iron sulfate'
Surfactant kg/t	0.0
Leach Temp °C	220
Sample	Blend ESPAN 3/CMT/SO
Heat up phase mins	49
Oxygen pressure Bar	6
Autoclave leach mins	60
Grind Size microns	108
Start Solution	
Volume ml	2000
Solution sg	1.08
Cu g/l	4
Fe (tot) g/l	12.98
Fe2+ g/l	12
Fe3+ g/l	0.98
As g/l	0.0068
Co ppm	3
H2SO4 g/l	23.39
Ferric/Ferrous ratio	0.08
Total SO ₄ ²⁻ g/l (calc)	52.20
End Solution	
Volume ml	-
Solution sg	-
Cu g/l	10.9
Fe (tot) g/l	36.1
Fe2+ g/l	-
Fe3+ g/l	36.1
As g/l	0.222
Co ppm	49
H2SO4 g/l	47.0
Total SO ₄ ²⁻ g/l (calc)	155.40
Metal Recovery % (based on solids)	
Weight loss %	
Cu	
Fe	
As	

Co	
Oxygen consumption g	163.0
Oxygen consumption kg/t	815.0
g Oxygen per g Cu leached	

Test No	AC409
Leach feed solids assay	
Weight gm	200
%Cu	6.91
%Fe	29.1
%As	0.24
Co ppm	640
S(sulphide) %	33.2
S(total) %	43.2
S(elem) %	9.03
S(sulphate) %	0.95
Leach residue solids assay	
Weight gm	na
%Cu	0.09
%Fe	29.6
%As	0.01
Co ppm	16

- Table 7 and FIG. 4 show results for an atmospheric leach of residue produced from a
- 5 220°C, 60 minute, autoclave test, designated AC409. Atmospheric leach conditions were: AC409 residue + 175ml of 12g/l H₂SO₄, 5g/l Fe²⁺, 0.4g/l Cu, 95°C, leach time 300 minutes. Under these conditions ferric iron is readily dissolved from basic iron sulfate produced in the autoclave. The increase of the ferric content from 9.1 g/L to 35.3 g/L validates the controlled redissolution of basic ferric sulfate and the
- 10 production of a rich ferric sulfate solution.

Table 7: Atmospheric re-leach of basic iron sulfate
 AC409 residue + 175ml of 12g/l H₂SO₄, 5g/l Fe²⁺, 0.4g/l Cu, 95°C

5

Time (minutes)	Fe ²⁺ g/l	Fe ³⁺ g/l	Fe _{total} g/l
0	0.6	9.1	9.7
60	0.7	29.6	30.3
120	1.3	33.7	35.0
180	1.4	33.9	35.3
300	1.1	35.3	36.4
end liquor SG:		1.17	

Example 3.

10 **Table 18: Summary of leach results – 220°C Tests with atmospheric post leach**

Test No	AC410	AC410 + atm
Description 3	220°C leach with atmospheric post leach with fresh feed to atm leach	treated as 1 test
Surfactant kg/l	0.0	0.0
Leach Temp °C	220	220
Sample	Sepon	Sepon
Heat up phase mins	52	52
Oxygen pressure Bar	6	6
Autoclave leach mins	30	30
Grind Size microns	57	57
Start Solution		
Volume ml	2134	2134
Solution sg	1.08	1.08
Cu g/l	4	4

Fe (tot) g/l	12.98	12.98
Fe2+ g/l	12	12
Fe3+ g/l	0.98	0.98
As g/l	0.0068	0.0068
Co ppm	3	3
H2SO4 g/l	23.39	23.39
Ferric/Ferrous ratio	0.08	0.08
Total SO ₄ ²⁻ g/l (calc)	52.20	52.20
End Solution		
Volume ml	-	
Solution sg	-	
Cu g/l	9.5	
Fe (tot) g/l	25.7	
Fe2+ g/l	-	
Fe3+ g/l	25.7	
As g/l	0.222	
Co ppm	49	
H2SO4 g/l	46.9	
Total SO ₄ ²⁻ g/l (calc)	126.61	
Weight loss %		35.4
Cu		87.9
Fe		46.6
As		57.9
Co		82.4
Oxygen consumption g	122.0	122.0
Oxygen consumption kg/l	610.0	366.4
g Oxygen per g Cu leached		8.45

Test No	AC410	AC410 + atm
Leach feed solids assay		
Weight gm	200	333
%Cu	4.93	4.93
%Fe	24.34	24.34
%As	0.46	0.46
Co ppm	55	55
S(sulphide) %	25.5	25.5
S(total) %	31.6	31.6
S(elem) %	6.05	6.05
S(sulphate) %	0.02	0.02
Leach residue solids assay		
Weight gm		215.21

%Cu		0.92
%Fe		20.1
%As		0.30
Co ppm		15

**Table 8: 220°C autoclave leach with partial bypass
and atmospheric leach of A/C residue**

5

Test No	AC410 + Atmospheric post leach
A/C Leach Temp °C	220
Sample	Sepon
Autoclave leach mins	30
Atmospheric Leach Temp °C	95
Atmospheric leach mins	360
Grind Size microns	57
Overall Metal Recovery % (based on solids)	
Weight loss %	35.4
Cu	87.9
Fe	46.6
Oxygen consumption kg/t	366.4
g Oxygen per g Cu leached	8.45

**Table 9: Atmospheric re-leach of AC410 residue + fresh feed
AC410 residue + 133g Sepon A/C feed + 200ml of
12g/l H₂SO₄, 5g/l Fe²⁺, 0.4g/l Cu, 95°C**

Time	Fe2+	Fe3+	Fe total
0	0.3	10.1	10.4
60	8.9	18.2	27.1
120	8.8	18.8	27.6
180	10.1	17.9	28.0
240	10.4	17.6	28.0
360	11.8	-	-
residue wt		215.21	
End liquor SG		.1156	

10

The results for this test show (1) high degree of iron dissolution in the atmospheric release and (2) excellent overall copper extraction.

It will of course be realised that while the above has been given by way of
5 illustrative example of this invention, all such and other modifications and variations thereto as would be apparent to persons skilled in the art are deemed to fall within the broad scope and ambit of this invention as is herein set forth.

DATED THIS THIRD DAY OF NOVEMBER, 2003.

10

OXIANA RESOURCES NL

by

PIZZEYS

Claim

1. A method for leach extraction of mineral bearing ores and concentrates including the step of leaching said mineral with an aqueous stream containing ferric ions and sulfuric acid in the presence of oxygen, at least part of said aqueous stream
- 5 comprising a solution formed by reaction of basic ferric sulfate with excess sulfuric acid.

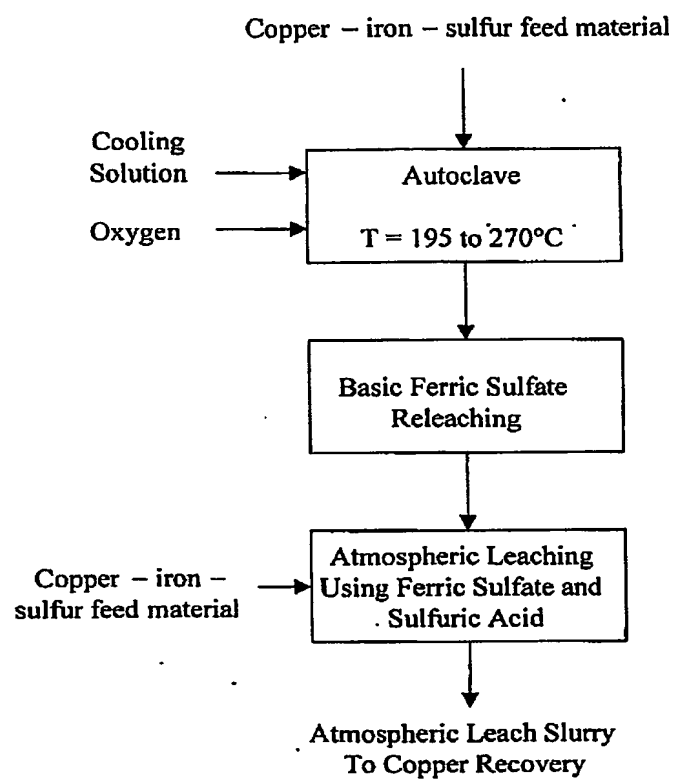


FIG. 1

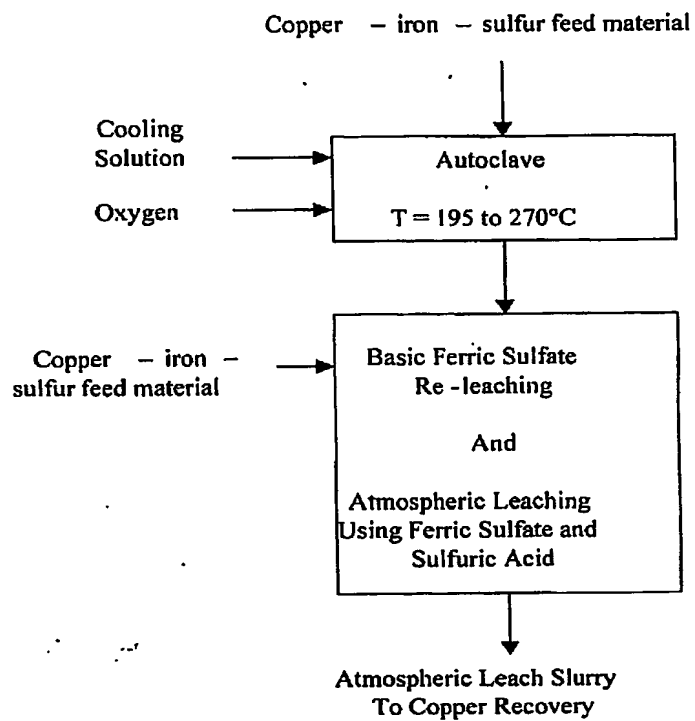


FIG. 2

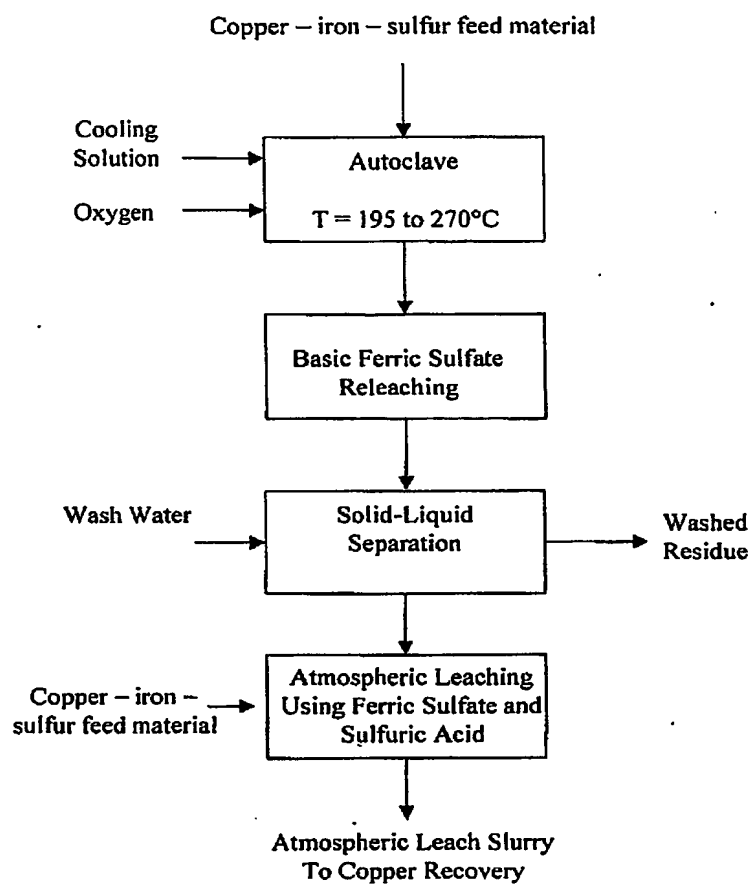


FIG. 3

Atmospheric re-leach of Basic iron sulfate
(Atmospheric re-leach of AC409 leach residue, 95 °C,
175ml liquor: 12g/l H_2SO_4 , 5g/l Fe^{2+} , 0.4g/l Cu)

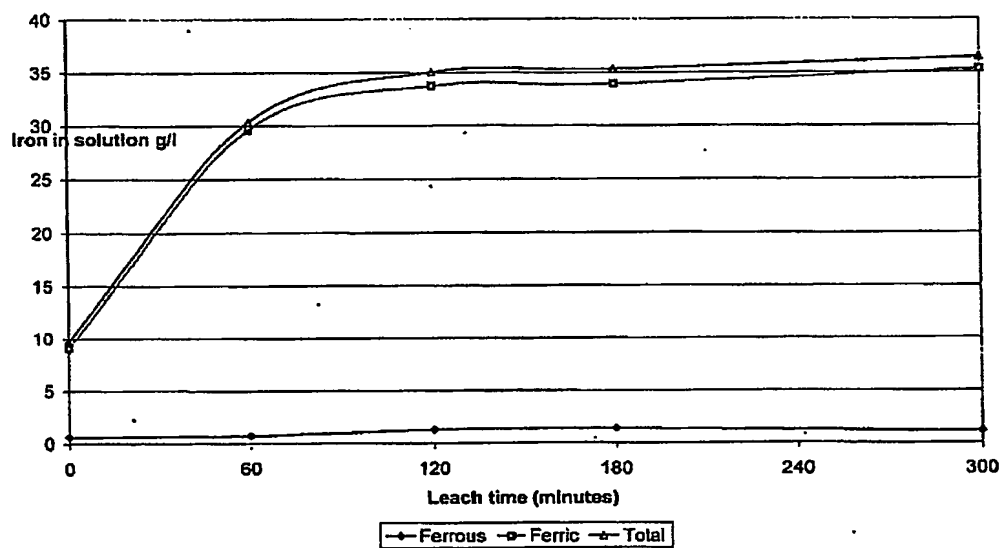


FIG. 4

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